**Chemistry - Equilibrium HW1 - Kinetics**

**Name: _____ Date: _____ Per: _____**

1. The _____ model explains why a reaction proceeds faster if the concentrations of the reacting molecules are increased.

2. True or false? A catalyst is a substance that speeds up a reaction without being consumed.
   A) True B) False

3. True or false? A minimum energy called the activation energy is needed for a reaction to occur.
   A) True B) False

4. True or false? At equilibrium, the concentrations of all reactants and products are constant.
   A) True B) False

5. True or false? At equilibrium, the concentrations of all reactants and products are equal.
   A) True B) False

6. Collisions need to have _____ and _____ for a reaction to occur.

7. Reaction rates can measure the disappearance of reactants and are stoichiometrically proportional.
   A) True B) False

8. All of the following are factors affecting the rate of a chemical reaction except _____.
   A) temperature
   B) surface area
   C) presence of a catalyst
   D) concentration
   E) volume

9. Increasing the surface area _____ the rate of a chemical reaction because it _____.

10. Reaction rates depend on factors such as _____, _____, and _____.

Bank: pressure, increases, collision, temperature, concentration