Name: _____ Mr. Leal
Date: _____ Per: _____ Chem/Honors - 11/8/2013

Molar / Formula Mass WS

A Mole:

Definition:

_____
_____

Molar Mass:

Definition:

_____
_____

The units for Molar Mass are: _____

Part 1: Write the Molar Mass for simple elements by locating the atomic mass on the Periodic Table:

- 1 mol of Carbon: _____
- 1 mol of Magnesium: _____
- 1 mol of Sodium: _____
- 1 mol of Tungsten: _____

Part 2: Calculate the Molar Mass (also known as Formula Mass) for compounds and molecules:

- **Determine the Molar Mass of Copper(II) Chloride:**

- **What I am given: _____**
- **Solve For: _____**

- **First determine the formula: _____**
- **Then, find the atomic masses for each element on the Periodic Table:**
  - **1 part Cu = 63.55 g**
  - **2 parts Cl = 35.45 g x 2 = 70.90 g**

- **Finally, add them up:**
  - **63.55 g + 70.90 g = 134.45 _____ (don't forget the units and formula!!)**

Your Turn!

Determine the molar mass for each of the following:

1. FeO

2. NaCl

3. H_2O

4. $BaSO_4$

5. $Mg(ClO_3)_2$ (hint: the number of atoms in the parenthesis are affected by the subscript 2 so there are 6 oxygens!)

6. Silver Nitrate

7. Ammonium Phosphate (remember what the parenthesis mean??)

8. Tin(IV) Permanganate

9. Potassium Acetate

10. Calcium Carbonate

11. Strontium Iodate

12. Triphosphorus Pentabromide