

Name: _____
Date: _____ Per: _____

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Chemistry - 11/8/2013

Moles WS 1 – Moles-Atoms

Converting Moles to Atoms

- To convert moles to atoms, multiply given by **Avogadro's Number** (_____ / **1 mole**)

$$\frac{1.50 \text{ mol C}}{1} = 9.03 \times 10^{23} \text{ atoms C}$$

Convert the following moles into atoms:

- 1.5 moles Al

- 0.535 moles Br

- 10.2 moles Au

Converting Atoms to Moles

- To convert atoms to moles, multiply given by **1 mole / Avogadro's Number**

$$\frac{5.0 \times 10^{23} \text{ atoms C}}{6.02 \times 10^{23}} = 0.83 \text{ mol C}$$

Convert the following atoms into moles:

- 1.5×10^{23} atoms C

- 2.5×10^{23} atoms K

- 222.5×10^{24} atoms Li