

Name: _____
Date: _____ Per: _____

Mr. Leal
Chemistry - 11/8/2013

Moles WS 3 – Conversions with Formulas

Converting Moles into Grams

- To convert moles into grams, multiply moles given by the molar mass:

$$\frac{2.0 \text{ mol C}}{1} = 24 \text{ g C}$$

Convert the following Moles into Grams:

- 2.5 moles Calcium Hydroxide
- 0.5 moles Iron (III) Chloride
- 1.55 moles Potassium Carbonate
- 0.000475 moles Lead (II) Iodide
- 1.039 mole Butane
- 2.023 moles Ammonium Nitrate
- 56.23 moles Copper(II) Sulfate
- 1.05×10^3 moles Hydrogen Sulfide
- 4.55×10^{-2} moles Barium Phosphate

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Converting Grams into Moles

- To convert grams into moles, multiply grams given by 1 mole / molar mass:

$$5.0 \text{ g C} = 0.42 \text{ mol C}$$

Convert the following Grams into Moles:

- 10.0 grams Carbon Dioxide
- 18.02 grams Copper (I) Bromide
- 32.91 grams Sodium Phosphate
- 834.091 grams Magnesium Nitrate
- 0.002378 grams Tetraphosphorus Pentabromide
- 48.01 grams Iron(II) Phosphide
- 39.190 grams Sucrose
- 2.25×10^4 mg Carbon Tetrafluoride