

## Polyatomic Compounds WS

Name \_\_\_\_\_

Chemistry

Per \_\_\_\_\_ Date \_\_\_\_\_

1. Add the proper charge for the simple ions and polyatomic ions listed. Example,  $\text{NH}_4$  becomes  $\text{NH}_4^+$ .
2. Add any ion symbols missing using the ion names provided.
3. Write the chemical formula for the combined pairs in the intersecting box. Example  $(\text{NH}_4)_2\text{SO}_4$

Anion → Cation ↓	OH	$\text{CO}_3$	$\text{PO}_4$	$\text{NO}_3$	$\text{SO}_4^{2-}$	$\text{Cr}_2\text{O}_7$
$\text{NH}_4^+$	1		2		$(\text{NH}_4)_2\text{SO}_4$	3
Iron(II)		4		5		
Tin(IV)			6		7	
Al		8		9	10	
K	11		12	13		14
Mg		15			16	17
Tin (II)	18	19		20		21
Iron (III)	22		23		24	