

1. Add the proper charge for the ions in the first row.
2. Fill in the missing ion names and ionic symbols for transition metal ions.
3. Write the chemical formula for the combined pairs in the intersecting box. Example  $\text{PbCl}_2$ .

Metal Ion Name	Metal Ion Symbol	F	$\text{Cl}^-$	O	Br	N	S	P
Tin(II)					1		2	3
	$\text{Pb}^{2+}$	4	$\text{PbCl}_2$	5		6		
Copper(I)			7			8	9	
	$\text{Fe}^{3+}$			10	11			12
Lead(IV)		13				14	15	
	$\text{Sn}^{4+}$		16			17	18	
Iron(II)				19	20			21
	$\text{Co}^{3+}$	22		23				24