

Name: _____

Date: _____ Per: _____

Mr. Leal

Chemistry - 11/14/2012

Stoich WS 1 – Formula-Eqn-Mole Ratios

To solve Stoichiometry problems, you may need to create equations from formulas and then balance them. This will allow you to create Mole Ratios. Mole ratios are setup as **Fractions** just like conversions.

Please create equations for the following reactions. Balance them, and then use them to create 4 different Mole Ratios:

1. Lead(II) Nitrate reacts with Sodium Iodide.
2. The combustion of C_2H_6 .
3. The reaction of phosphorus trichloride with water to create phosphorous acid and hydrochloric acid.
4. The single displacement of Iron(III) Oxide with Aluminum.
5. The decomposition of aqueous hydrogen peroxide into a liquid and a gas.

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Using Mole Ratios to solve Stoich problems:

Using the same equations, we can solve for Moles if we are given any number of moles. For each equation, we will start with 1.50 moles of the first reactant. Determine the number of moles for all other formulas.

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2. The combustion of C_2H_6 .
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5. The decomposition of aqueous hydrogen peroxide into a liquid and a gas.